

SUBJECT TEACHING GUIDE

G986 - Chemistry

Degree in Industrial Electronic Engineering and Automatic Control Systems

First Degree in Industrial Electronic Engineering and Automatic Control Systems

Academic year 2024-2025

1. IDENTIFYING DATA					
Degree	Degree in Industrial Electronic Engineering and Automatic Control Systems First Degree in Industrial Electronic Engineering and Automatic Control Systems			Type and Year	Core. Year 2 Core. Year 1
Faculty	School of Industrial Engineering and Telecommunications				
Discipline	Subject Area: Chemistry Basic Training Module				
Course unit title and code	G986 - Chemistry				
Number of ECTS credits allocated	6	Term	Semester based (1)		
Knowledge Field					
Web					
Language of instruction	Spanish	English Friendly	No	Mode of delivery	Face-to-face

Department	DPTO. DE QUIMICA E INGENIERIA DE PROCESOS Y RECURSOS.				
Name of lecturer	GEMA RUIZ GUTIERREZ				
E-mail	gema.ruiz@unican.es				
Office	E.T.S. de Ingenieros Industriales y de Telecomunicación. Planta: - 3. DESPACHO (S3087)				
Other lecturers	LORENA GONZALEZ LEGARRETA				

4. OBJECTIVES

This subject's goal is to provide students with basic knowledge on chemistry to develop chemistry related skills in relation with their future professional performance as engineers

6. SUBJECT PROGRAM	
CONTENTS	
1	<p>MODULE 1. BASIC PRINCIPLES OF CHEMISTRY AND THEIR APPLICATION TO ENGINEERING SYSTEMS.</p> <p>Lesson 1. ATOMS, ELEMENTS AND COMPOUNDS. Origins of modern chemistry. Taxonomy of chemical species. Introduction to the periodic table. Periodic relationships among the elements. Reactions in aqueous solutions.</p> <p>Lesson 2. INTRODUCTION TO BONDING. Ionic bonding, lattice energy. Properties of ionic crystal. Covalent bond, polar bonds and polar molecules. Intermolecular forces. Properties of covalent molecules. Properties of metals, semiconductors, and insulators.</p> <p>Lesson 3. THERMOCHEMISTRY. Energy changes in chemical reactions. Introduction to thermodynamics. Enthalpy in chemical reaction. Calorimetry. Standard enthalpy of formation and reaction. Heat of solution and dilution.</p> <p>PRACTICAL CLASSES LABORATORY SESSIONS 1, 2 AND 3</p>
2	<p>MODULE 2. CHEMICAL EQUILIBRIUM AND ELECTROCHEMISTRY.</p> <p>Lesson 4. CHEMICAL EQUILIBRIUM. Chemical equilibrium, equilibrium constant, heterogeneous equilibria, multi-stage equilibria, degree of dissociation, factors that modify the chemical equilibrium.</p> <p>Lesson 5. ACID-BASE EQUILIBRIUM. Properties of acids and bases, definitions of acids and bases, relative strengths of acids and bases, self-ionization of water, pH scale, ionization constant of acid and base, ratio between ionization constants acid-base conjugate pairs, diprotic and polyprotic acids, acid-base properties of salts, buffer solutions.</p> <p>Lesson 6. OXIDATION-REDUCTION REACTIONS. Redox reactions, oxidation number, balancing redox reactions.</p> <p>Lesson 7. ELECTROCHEMICAL CELL. Electrochemical cell, standard electrode potential, energy of electrochemical cell, effect of concentration on cell emf, Nernst equation, concentration cell.</p> <p>Lesson 8. ELECTROLYTIC CELL. Electrolysis, quantitative aspects of electrolysis.</p> <p>PRACTICAL CLASSES LABORATORY SESSIONS 4 AND 5</p>

7. ASSESSMENT METHODS AND CRITERIA				
Description	Type	Final Eval.	Reassessn	%
Exam part 1	Written exam	No	Yes	35,00
Exam part 2	Written exam	Yes	Yes	35,00
Laboratory classes	Laboratory evaluation	No	Yes	15,00
Personal work and practical classes (problem solving)	Work	No	No	15,00
TOTAL				100,00
Observations				
<p>To pass by continuous assessment, the four parts will be taken into account: the delivery of the exercises and problems collected during the classes, the laboratory practices and the evaluation tests 1 and 2. The final grades of the ordinary and extraordinary calls will be determined considering the following percentages: 15% the delivery of exercises and/or problems, 15% the laboratory practices, 35% evaluation test 1 and 35% evaluation test 2.</p>				
Observations for part-time students				
<p>For part-time students, single assessment. The students will be able to pass the subject by presenting themselves to the evaluation exams 1 and 2, and/or the extraordinary exam.</p>				

8. BIBLIOGRAPHY AND TEACHING MATERIALS
BASIC
American Chemical Society, "Química" un proyecto de la ACS. Ed. Reverté, 2007.
R. Chang, "Química". Ed. Mc Graw Hill, (9ª Ed.) 2007.
R.H. Petrucci, W.S. Harwood, F.G. Herring. "Química General". Ed. Mc Graw Hill, 2007.
J. Vale, C. Fernández, M. Piñero, M. Alcalde, R. Villegas, L. Vilches, B. Navarrete, Á García. "Problemas resueltos de Química para Ingenieros." Ed. Thomson, 2004.